

I'm not robot!

Types of Sentences

There are four types of sentences in the English language:

- **Declarative:** this sentence makes a statement. Use a period (.) to punctuate a declarative sentence.
Example: Dogs require daily walks.
- **Interrogative:** this sentence asks a question. Use a question mark (?) to punctuate an interrogative sentence.
Example: Did you know that dogs require daily walks?
- **Imperative:** this sentence gives a command (an order). Use either a period (.) or an exclamation mark (!) to punctuate an imperative sentence. The difference is one of tone (if you are expressing emotion, like anger, use an exclamation mark).
Examples: Walk the dog. Go walk the dog already!
- **Exclamatory:** this sentence expresses strong emotion, like pain, anger, joy. Use an exclamation mark (!) to punctuate an exclamatory sentence.
Example: I got a dog yesterday!

A. Identify the type of sentence for all of the sentences below. Write **D** for declarative, **I** for interrogative, **IM** for imperative, or **E** for exclamatory before each sentence.

1. How did you get to school today?
2. My mother bakes the best bread.
3. It is so cold today!
4. There are many books in Miss Rich's class.
5. Go to your room!
6. Did you know that a moose is taller than most cars?
7. I love this song!
8. I think my leg is broken!
9. Would you like to go to Ireland with me?
10. Close the door when you leave.

LIVEWORKSHEETS



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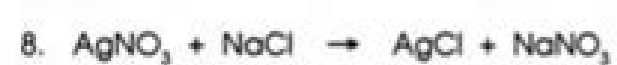
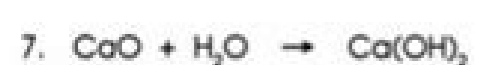
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Contributing to Each Other's Learning

CLASSIFICATION OF CHEMICAL REACTIONS

Name _____

Classify the reactions below as synthesis, decomposition, single replacement (cationic or anionic) or double replacement.



Chemistry #8766

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Introduction to Chemical Reactions

- ✓ Engagement
- ✓ Exploration
- ✓ Explanation
- ✓ Elaborate
- ✓ Evaluate



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WORD EQUATIONS Name _____

Write and balance the following chemical equations.

- Hydrogen plus oxygen yields water. $2H_2 + O_2 \rightarrow 2H_2O$
- Nitrogen plus hydrogen yields ammonia. $N_2 + 3H_2 \rightarrow 2NH_3$
- Aluminum bromide plus chlorine yields aluminum chloride and bromine. $2AlBr_3 + 3Cl_2 \rightarrow 2AlCl_3 + 3Br_2$
- Hydrochloric acid plus sodium hydroxide yields sodium chloride plus water. $HCl + NaOH \rightarrow NaCl + H_2O$
- Iron plus lead (II) sulfate react forming iron (II) sulfate plus lead. $Fe + PbSO_4 \rightarrow FeSO_4 + Pb$
- Potassium chlorate when heated produces potassium chloride plus oxygen gas. $2KClO_3 \rightarrow 2KCl + 3O_2$
- Sulfuric acid decomposes to form sulfur dioxide gas plus water. $H_2SO_4 \rightarrow SO_2 + H_2O$
- Sodium oxide combines with water to make sodium hydroxide. $Na_2O + H_2O \rightarrow 2NaOH$
- Potassium iodide reacts with bromine forming potassium bromide plus iodine. $2KI + Br_2 \rightarrow 2KBr + I_2$
- Sodium phosphate reacts with calcium nitrate to produce sodium nitrate plus calcium phosphate. $2Na_3PO_4 + 3Ca(NO_3)_2 \rightarrow 6NaNO_3 + Ca_3(PO_4)_2$
- Zinc reacts with iron(III) chloride yielding zinc chloride plus iron precipitate. $3Zn + 2FeCl_3 \rightarrow 3ZnCl_2 + 2Fe$
- Ammonium carbonate and magnesium sulfate react to yield ammonium sulfate plus magnesium carbonate. $(NH_4)_2CO_3 + MgSO_4 \rightarrow (NH_4)_2SO_4 + MgCO_3$
- Phosphoric acid plus calcium hydroxide react forming solid calcium phosphate plus water. $2H_3PO_4 + 3Ca(OH)_2 \rightarrow Ca_3(PO_4)_2 + 6H_2O$
- Aluminum plus oxygen gas form aluminum oxide under certain conditions. $4Al + 3O_2 \rightarrow 2Al_2O_3$
- Nitrogen gas plus oxygen gas react and form dinitrogen pentoxide. $2N_2 + 5O_2 \rightarrow 2N_2O_5$

Classifying chemical reactions worksheet answers instructional fair.

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